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spectrum of the gas mixture is determined. The absorption spectrum of the gas mixture is typically a function of its concentration and of its molecular structure. In addition, the spectra of the individual components may differ. Because of this and because the absorption characteristics of the gas mixture and the gas components may vary under changing environmental conditions, the analysis of the gas mixture often proceeds in a plurality of steps or cycles. For example, a sample gas mixture is irradiated with infrared radiation in a first step. In a second step, a specific absorption band is excited with a radiation source. The excitation may result in a fluorescence of the sample. In a third step, the sample is passed through a filter that transmits only the fluorescence. The filtered radiation is converted with a detector into an electrical signal. The electrical signal is then analyzed in a fourth step. Because the measured values of the spectral absorption and the spectral fluorescence values are usually complex functions of the gas distribution of the sample, infrared gas analysis methods typically include a mathematical simulation or a correction of the measured values by a calibration. For example, U.S. Pat. No. 6,239,562 discloses an infrared gas analysis method for an oxygen concentration measurement in which the oxygen concentration is deduced from the measurement of f988f36e3a

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